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~~Ions How to Memorize The~~

~~Polyatomic Ions - Formulas,~~

~~Charges, Naming - Chemistry~~

~~*Naming Compounds in*~~

~~*Chemistry How to calculate*~~

~~*valency? VSEPR Theory:*~~

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11: Section 1: Describing
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Formulas

How to Write Chemical
Formulas from Compound Names

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Governing How Compounds are
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~~Chemical name and formula of
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Valence Electrons and the
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Formula and Molecular
Formula Introduction

*Covalent Compounds: Writing
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~~Chapter 9: Chemical Names~~

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Chapter 9: Chemical Names
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Monatomic Ion. Consists of a
single atom with a positive
or negative charge resulting
from the loss of gain of one

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or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula
2. Use prefixes appropriately to indicate

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the number of each kind of
atom 3. End the name of the
second element with the
suffix -ide

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Names and Formulas
Flashcards ...~~

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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the

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~~Formula Answers~~ only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric

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acid) 2. when the anion name ends in -ite the acid name is the stem of the anion with the suffix -ous followed by the word acid (ex. H_2SO_3 -- X=sulfite = sulfurous acid)

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~~CHAPTER 9 Chemical Names And
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Writing the Names and
Formulas for Bases. A base
is a compound that produces
 OH^- in water. When naming a
base, you name it like any
other compound that starts

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with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the formula for a base, remember to balance charges. Ex: magnesium hydroxide = $\text{Mg}(\text{OH})_2$

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Chapter 9 Chemical Names and
Formulas 83 SECTION 9.3
NAMING AND WRITING FORMULAS

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FOR MOLECULAR COMPOUNDS

(pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the

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Letter of the type(s) of elements that form binary molecular compounds.

~~Name Date Class CHEMICAL
NAMES AND FORMULAS 9
SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC~~

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COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium

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sulfide g. CuCl_2 h. SnCl_2

a. MnO_2 c. d. SrBr_2 221

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281 CHAPTER 9 Assessment 42.

a. 2- b. 1+ c. 1- d. 3+ 43.

a. 2+ b. 2+ c. 3+ d. 1+ 44.

a. barium ion b. iodide ion

c. silver ion d. mercury(II)

ion 45. cyanide, CN^- and

hydroxide, OH^- 46. a.

hydroxide ion b. lead(IV)

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ion c. sulfate ion d. oxide
ion 47. zero; A compound is
electrically neutral. 48.
The symbols for the cation
and

~~CHAPTER 9 Study Guide — Quia~~

Chapter 9 Chemical Names

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STUDY. Flashcards. Learn.
Write. Spell. Test. PLAY.
Match. Gravity. Created by.
bibiloni1901. Terms in this
set (20) Monatomic Ion.
Consists of a single atom

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with a positive or negative charge resulting from the loss of gain of one or more valence

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~~Packet Answers~~ Chapter 9; anion: any atom or group of atoms that has a negative charge:

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Formulas Answer Section

MATCHING 1. ANS: A PTS: 1

DIF: L1 REF: p. 253 OBJ:

9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS:

H PTS: 1 DIF: L1 REF: p. 254

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