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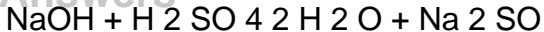
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following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$

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3 Using the following equation: 2



4

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Review Worksheet. 1) Using the
following equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$ How many
grams of sodium sulfate will be formed
if you start with

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Answers

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following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2$

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NF₃ 2) 2 C₆H₁₀ + 17 O₂ 12 CO₂
+ 10 H₂O 3) 1 HBr + 1 KHCO₃ 1 H₂
O + 1 KBr + 1 CO₂ 4) 2 GaBr₃ + 3
Na₂SO₃ ... and the reagent that
leads to this answer is the limiting
reagent. Both Balancing Equations
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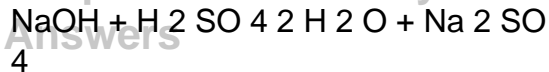
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2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$
5) $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}$

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3 Using the following equation: 2



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The first step to balance the equation is to write down the chemical formula

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of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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~~49 Balancing Chemical Equations
Worksheets [with Answers]~~

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following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$

2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2$

+ $10 \text{ H}_2\text{O}$ 3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O}$

+ $1 \text{ KBr} + 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3$

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Na₂SO₃ + Ga₂(SO₃)₃ + 6 NaBr

5) 3 SnO + 2 NF₃ → 3 SnF₂ + 1 N₂O

3 Using the following

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Review Worksheet. 1) Using the

following equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4$

$\rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$ How many

grams of sodium sulfate will be formed

if you start with 200 grams of sodium

hydroxide and you have an excess of

sulfuric acid (H_2SO_4)? 2) Using the

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following equation: $\text{Pb}(\text{SO}_4)$

Answers

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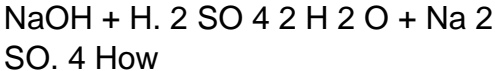
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1) Using the following equation: 2



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